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The vibrating molecules emitting the radiation could have only certain discrete amounts of energy, given by $E_n = nhf$, where n is called a quantum number, f is the frequency, and h is Planck's constant given by $h = 6.626 \times 10^{-34}$ joule·seconds. Molecules emit energy in units called quanta, now called photons.

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$Z \cdot \nabla \cdot \psi(r,t) = 0$. (6) This result is proven to be in agreement with the Schrödinger equation, and it then becomes natural to define the probability density ρ and probability current density j according to $\rho(r,t) = |\psi(r,t)|^2$, $j(r,t) = \text{Re}$.

Study Guide for Quantum Mechanics (TFFY54)

Quantum mechanics is the branch of physics dealing with physical phenomena at microscopic scales, particularly with notions of probability. Probability in Matter. Considering light as particles (photons), the probability per volume of finding a photon in a given region of space at a given time is proportional to the number N of photons per unit volume at that time and to the intensity:

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