

## Pogil Chemistry Answer Key Equilibrium

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Consider the following reaction  $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$  At 425°C, an equilibrium mixture has the following concentrations.  $[\text{HI}] = 1.01 \times 10^{-2} \text{ mol/L}$ .  $[\text{H}_2] = 1.25 \times 10^{-3} \text{ mol/L}$ .  $[\text{I}_2] = 1.49 \times 10^{-3} \text{ mol/L}$ .

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### Equilibrium Continued - Solutions

Equilibrium in chemical reactions is dynamic because the forward and reverse reactions are occurring continuously and simultaneously at the same rates. Placing a stress on any equilibrium system, whether it is chemical, biological, societal, environmental, or personal, causes the equilibrium position to change.

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### Mrs. Zuberbuehler - Mrs. Zuberbuehler

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- Chemical equilibrium occurs when opposing reactions are proceeding at equal rates. • The rate at which the products are formed from the reactants equals the rate at which the reactants are formed from the products.

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### CHEMICAL EQUILIBRIUM (ICE METHOD)

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### Buffers - SUPERTALLTEACHER

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Pogil Biology Answer Key Cell Size - [cdnx.truyenyy.com](http://cdnx.truyenyy.com) Download Ebook Dynamic Equilibrium Pogil Answers at equilibrium:  $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{H}_2\text{O}(\text{g}) + \text{energy (heat)}$  In Table 1 apply Le Chatelier's Principle and indicate the direction of the shift in equilibrium if the indicated stress is applied to the reaction system.

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## Instructors Guide Pogil Equilibrium Answer Key

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