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following  
statements,  
write "I" for  
ionic, "C" for  
covalent

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unit for the  
following  
measurements.

Define

temperature and

mass. a. Length

b. mass c.

volume d.

temperature e.

time f. quantity

9. Perform the

following unit

conversions: Use

the factor label

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year of high  
school chemistry  
or the first  
semester of

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the terms, units  
of measurement,  
and atoms,  
molecules,  
elements, and  
compounds may  
seem a bit  
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course is taught  
as a  
continuation of

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the first year  
course. As such,  
units are  
numbered

accordingly.

Units 1-8 are  
covered during  
the first year  
course and  
reviewed over a  
short time span  
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Part 2. Read the descriptions below and

determine

whether each

item describes

an endothermic

or an exothermic

reaction. \_\_\_\_\_

1. A reaction

that produces

heat energy

\_\_\_\_\_ 2. A

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reaction that  
absorbs energy  
\_\_\_\_\_ 3. Paper  
burns \_\_\_\_\_ 4

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Show all  
questions <= ...  
2.5 moles per  
second; Given

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the following  
reaction:  $C_3H_8(g) + 5O_2(g) \rightarrow 3CO_2 + 4H_2O(g)$  If the  
initial rate of  
the reaction  
produces 2 moles  
of water vapor,  
 $H_2O$ , per  
second, what is  
the rate at  
which propane,  $C_3H_8$ , is

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initially? ?  
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36.5 g H M =  
moles of solute  
liters of  
solution 0.200 M  
= mol solute  
0.250 L mol  
solute = 0.0500

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mol HCl? g HCl =

0.0500 m HCl !

Cl 1 mol HCl =

Cl 4. A solution

of Na<sub>2</sub>CO<sub>3</sub>

contains 65.0 g

of solute

dissolved in

water to make a

3.00 M solution.

What is the

volume of the

solution, in

liters? ? g HCl

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$= 0.0500 \text{ mol HCl}$   
 $! 36.5 \text{ g HCl} \cdot 1$   
 $\text{mol HCl} = 1.83$

10-24a, b, c, d

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Chemistry 12

Unit 1 -

Reaction

Kinetics Unit 1

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each of the  
following  
reactions find a  
quantity or  
property which  
could be  
monitored in  
order to measure  
the rate of  
reaction See p.  
2-5 in SW. ("a"  
is done as an  
example.) a)  
 $3\text{H}_2(\text{g}) + \text{N}_2(\text{g}) \rightarrow$

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$2\text{NH}_3(\text{g})$  pressure  
will decrease as  
reaction  
proceeds because

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