

Chemistry Post Lab Answers

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~~CHEM 131 L- Experient 9 - The questions and answers for ...~~

Answers to Post-Lab Questions for Experiment 5. I have 40 L of methane gas measured at standard temperature and pressure (STP). How many moles of methane gas is this? 1 mole of any gas at STP = 22.4 L Moles of methane = $(40 / 22.4) = 1.79$ moles. What is the volume of 0.30 moles of nitrogen gas if measured at STP? 1 mole of any gas at STP = 22.4 L Volume of nitrogen = $0.3 \times 22.4 = 6.72$ L.

~~SCG1123 Experiment 5 Post-Lab Questions Answers - ECU ...~~

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Favorite Answer. First calculate the moles of KHP (potassium hydrogen phthalate, $\text{KHC}_8\text{H}_4\text{O}_4$). The molar mass for KHP = 204.2 g/mole. $0.531 \text{ g KHP} \times (1 \text{ mole KHP} / 204.2 \text{ g KHP}) = 0.00260 \text{ moles KHP}$. KHP...

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Favorite Answer. The white precipitate most likely came from the presence of chloride ions. $\text{Ag}^+ + \text{Cl}^- \rightarrow \text{AgCl}(s)$ The chemical responsible for the iron is Fe (iron), in some form. It could be almost...

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Follow this series of questions. When you can answer " yes: " to a question, then stop. 1) Does your reaction have oxygen as one of it's reactants and carbon dioxide and water as products? Yes: it's a combustion reaction 2) Does your reaction have two (or more) chemicals combining to form one chemical? Yes: it's a synthesis reaction

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Chemistry is an experimental science. Thus, it is important that students of chemistry do experiments in the laboratory to more fully understand that the theories they study in lecture and in their textbook are developed from the critical evaluation of experimental data.

~~Chemistry 2A Lab Manual~~

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