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[2] (b) Marit arranged the values she found in Table 12 into an energy cycle.  $\text{C}_2\text{H}_4(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$   
 $302 \text{ kJ mol}^{-1}$   
 $2\text{CO}_2(\text{g}) + 3\text{H}_2\text{O} - 1411 \text{ kJ mol}^{-1}$   
 $-286 \text{ kJ mol}^{-1}$   
 $-1560 \text{ kJ mol}^{-1}$   
(hydrogenation) (l) Calculate the value for the enthalpy of hydrogenation of ethene from the energy cycle.....

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