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H ? mol C H = 5.5 mol

O = 0.85 mol C H 13

mol O ... 65XI0 Grams

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Practice Problems:

Moles (Answer Key)

How many moles are in the following: a. 1.29×10^{24} hydrogen atoms

in HF 2.14 moles H atoms.

b. 7.36×10^{24} free oxygen atoms 12.2 moles O atoms.

c. 3.28×10^{23} Na atoms in salt (NaCl) 0.545 moles Na

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atoms. How many
atoms are present in the
following?

Practice Problems:

Moles (Answer Key)

Reveal answerupdown.

M_r of NaOH = $23 + 16$

+ $1 = 40$. M_r of Na₂

SO₄ = $23 + 23 + 32 +$

$16 + 16 + 16 + 16 =$

142. Number of moles

of NaOH = mass ÷

relative formula mass =

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20 ÷ 40 = 0.5 mol. From

Key

Mole calculations -

Formula mass and mole
calculations ...

2, moles H₂ =

$(1.7)(3)/2 = 2.6$ moles b.

3 moles H₂SO₄ gives

1 mole product moles

yield = $3 \times 2.8 = 8.4$

moles 3. Mole ratios: 2

mol Mg/ 2 mol MgO =

1 mol Mg: 1 mol

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product 1 mol O_2 / 2
mol $MgO = 2$ mol O_2 :
1 mol product 4. Mole
ratios: 2 mol Al / mol Al
 $2O_3$, 1.5 mol O_2 / mol
 Al_2O_3 . Hence, 5 mol
 Al and 3.75 mol O_2 are
needed 5. Mole ratios: 1
mol $BaCl_2$: 1 mol
 $BaSO_4$

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Key Chapter 12 ...
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Mole Conversions

Chem Worksheet 11 3

Answers - worksheet

4.3×10^{22} molecules

$\text{H}_3\text{PO}_4 \times 1 \text{ mole}$

$\text{H}_3\text{PO}_4 = 7.1 \times 10^{-2}$

moles H_3PO_4 . 6.022×10^{23} molecules H_3PO_4 .

5) How many molecules are in 48.0 grams of

NaOH ? 48.0 molecules

$\text{NaOH} \times 1 \text{ mole NaOH}$

$\times 6.022 \times 10^{23}$

molecules NaOH . 40 g

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NaOH 1 mole NaOH. =
 7.23×10^{23} molecules
NaOH.

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We begin by converting
the mass of CH₄ to
moles of CH₄, using
the molar mass of CH₄
(16.05 g/mol) as the
conversion factor:

$\frac{100.0 \text{ g}}{16.05 \text{ g/mol}}$

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$\text{CH}_4 \times \frac{1 \text{ mol CH}_4}{16.05 \text{ g CH}_4} = 6.231 \text{ mol CH}_4$) Note that we inverted the molar mass so that the gram units cancel, giving us an answer in moles.

6.5: Mole-Mass and
Mass-Mass Problems -
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Using $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$, how many moles of

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hydrogen react with
3.07 mol of oxygen to
produce H₂O? Answer
$$\left(\frac{3.07 \text{ mol O}_2}{2} \right) \times \left(\frac{1 \text{ mol H}_2}{1 \text{ mol O}_2} \right) = 6.14 \text{ mol H}_2$$

6.4: Mole-Mole
Relationships in ... -
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The molar mass of
water is (1.0079 +

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$(1.0079 + 15.999) =$
18.015 g/mol. However,
because we want to
cancel the gram unit and
introduce moles, we
need to take the
reciprocal of this
quantity, or 1
mol/18.015 g: Test
Yourself. How many
moles are present in
35.6 g of H_2SO_4
(molar mass = 98.08
g/mol)? Answer. 0.363

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question on Mole
Calculations. More
specifically it involves
skills in balancing
chemical equations and
writing ionic equations.
PS: Try it out and leave
your suggested answer

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at the "Comments
Section" right below
this post. Question:

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with nitric acid to form
3 other products.□

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