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Chapter 1: Matter and

Change (Chem in 15

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1 (Chem 1 H) GCSE

Science Revision

Chemistry \"The Three

States of Matter\" The

Creation of Chemistry -

The Fundamental Laws:

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#3 Matter and Change

~~Chapter 1~~

~~Introduction: Matter and~~

~~Measurement States of~~

~~matter | States of matter~~

~~and intermolecular~~

~~forces | Chemistry |~~

~~Khan Academy Grade 9~~

~~Chemistry Lesson 1 -~~

~~Matter and the Particle~~

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Particulate nature of
matter - Part-1 -

Chemistry Changes of
States of Matter

Chemical and Physical
Changes

States of Matter : Solid
Liquid Gas 10 Amazing
Experiments with Water

Pure Substances and
Mixtures!

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(Classification of
Matter)Physical and
Chemical Changes

Understanding Atoms,
elements, and molecules
Part #1 (9min)

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1.2 What is matter? -
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Physical and Chemical
Changes: Chemistry for
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~~Changes States of
Matter Solids, Liquids,~~

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~~Chemistry Change of~~
~~State | Matter | Physics |~~
~~FuseSchool~~ States of
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Mixtures, Elements
\u0026 Compounds,
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Chemistry Examples,

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Introduction to

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Stories of Two

Chemicals Ozone Layer,
atmosphere, ozone

formation,

chlorofluorocarbons,

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Chapter 1.2 Chemistry
and Matter Chemistry
Central Science

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Chapter Summaries □

Chemistry Matter and
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Chapter 1 Section 2

MATTER and Its

Properties Matter □

Anything that has

MASS and VOLUME.

ATOMS and

MOLECULES the basic

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building blocks of matter. ATOM-- the smallest unit of an element that still maintains the properties that element.

ELEMENT □ a pure substance made up of only one type of atom.

MATTER and
CHANGE Chapter 1
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CHAPTER Class

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Section 6.3 Periodic

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atomic radius and ionic

radius. Circle the letter

of the choice that best

completes the statement

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1. Atomic radii cannot
be measured ...

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CHAPTER 3

SOLUTIONS

MANUAL

Matter Properties and C

hanges Matter Properties

and Changes Solutions

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Chapter 3 35 Section 3.1

Properties of Matter

pages 70□75 Problem-

Solving Lab 1. Explain

why the flow of a

compressed gas must be

controlled for practical

and safe use.

Matter□Properties and C

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Guide - Chapter 6 □ The
Periodic Table and
Periodic Law Section
6.1 Development of the
Modern Periodic Table
1. octaves 2. eight 3.
nine 4. accepted 5.
Dmitri Mendeleev 6.
atomic mass 7. elements

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8. atomic number 9.
Henry Moseley 10.
protons 11. periodic law
12. properties 13. 14 ...

Answers

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answers

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Chapter 5 25 Electrons
in Atoms Reviewing
Vocabulary Match the
definition in Column A

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with the term in Column

B. Column A Column B

1. The set of frequencies

of the electromagnetic

waves emitted by the

atoms of an element 2.

The minimum amount

of energy that can be

lost or gained by ...

Chapter Assessment

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of Matter - 2.1 Lesson
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ISBN-10: 0132525763,

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...

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numbers used in

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chemistry can be defined as the study of energy changes (that is, energy produced or absorbed).

A system that can exchange neither energy nor matter with its surroundings is called a closed system.

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extensive property.

intensive property.

extensive properties

(examples) intensive

property (examples) a

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property that depends on the amount of matter in a sample. a property that depends on the type of matter in a sample not. mass, volume, length, weight.

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number of moles of
known substance C H C

H C H O O O CO CO

CO H O H O H O

number of moles of

unknown substance O

CO H O C H CO H O C

H O H O C H O CO use

mole ratio 5 mol O 2 1

mol C 3 H 8 3 mol CO 2

1 mol C 3 H 8 4 mol H

2 O 1 mol C 3 H 8 1

mol C 3 H 8 5 mol O 2

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MASTER 16

A [chemical change] is a change that produces matter with a different composition than the original matter.

Chemical Change A change in which one or more substances are converted into different substances. Heat and

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light are often evidence
of a chemical change.

Properties of
Compounds Quite

different properties than
their component
elements.

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