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Ion. Consists of a

single atom with a

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positive or negative charge resulting from the loss of gain of one or more valence electrons and behaves as a unit, yet it still has a charge. (+ or -). Determined by the amount of electrons lost or gained.

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single atom with a

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gain of electrons,
respectively.

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Naming and Writing

Formulas for Acids

and Bases. An acid

is a compound that

produces H^+ ions

when it dissolves in

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water. The formula for an acid normally starts with and H.

When naming acids, you should first determine whether or not the acid contains oxygen.

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base. Cation. Ionic

compounds.

Compounds that

contain one or more

hydrogen atoms and

produce.... An ionic

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Compound that produces hydroxide ions when dissolved.... Any atom or group of atoms that has a positive charge.

Compounds composed of cations and anions.

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SECTION 9.3

NAMING AND
WRITING

FORMULAS FOR
MOLECULAR
COMPOUNDS

(pages 268 – 270)

This section
explains the rules
for naming and

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writing formulas for binary molecular compounds. Naming Binary Molecular Compounds (pages 268 – 269) 1. Circle the letter of the type(s) of elements that form binary molecular compounds.

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~~AND FORMULAS 9~~

SECTION 9.2 NAMES AND WRITING

FORMULAS FOR IONIC COMPOUNDS 1.

Write the formulas for these binary ionic compounds. c. potassium iodide 2. write the formulas for the compounds formed c. b. f. 3.

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Name the following binary ionic compounds. sodium sulfide g. CuCl : h. SnCl : a. MnO_2 c. d. SrBr_2 221 Chapter 9 Chemical Names and Formulas

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and Formulas ” ... in
parenthesis after
the name of the
metal (Table 9.2,
p.255) Predicting
Ionic Charges Some

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of the post-transition elements also have more than one possible oxidation state. Tin
(II) = Sn^{2+} Lead
(II) = Pb^{2+} Tin
(IV) = Sn^{4+} Lead
(IV) = Pb^{4+}

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2 – b. 1 + c. 1 – d. 3 +

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43. a. $2+$ b. $2+$ c.
 $3+$ d. $1+$ 44. a.
barium ion b. iodide
ion c. silver ion d.
mercury(II) ion 45.
cyanide, CN^- and
hydroxide, OH^- 46.
a. hydroxide ion b.
lead(IV) ion c.
sulfate ion d. oxide
ion 47. zero; A
compound is
electrically neutral.
48. The symbols

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for the cation and

Names And

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Chapter 9 Practice
Problems: Chemical
Names and

Formulas. Section
9.1: Naming Ions. 1.
What is the charge
on the ion typically
formed by each
element? a. oxygen
c. sodium e. nickel,

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two electrons lost .

b. iodine d.

aluminum f.

magnesium. 2. How

many electrons

does the neutral

atom gain or lose

when each ion

forms? a. Cr^{3+} c.

Li^+ e. Cl^- b. P^{3-} ...

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In samples of any chemical compound, the masses of the elements are always in the same proportions any atom or group of atoms that has a negative charge ... place cation name

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first followed by the
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